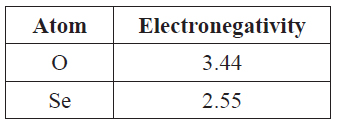
**Electronegativity**

**1)** Using your knowledge of trends in the periodic table, circle the atom below that has the greater electronegativity value. Justify your answer.

**Br I**

**2)** Discuss the data for the following



**3)** Account for the differences in electronegativity values for nitrogen, potassium and arsenic.

|  |  |
| --- | --- |
| **Element** | **Electronegativity** |
| Nitrogen, N | 3.04 |
| Potassium, K | 0.82 |
| Arsenic, As | 2.18 |

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